

Reactivity of Elements

- Just as a fingerprint is unique to each person, the color emitted by metals heated in a flame is unique to each metal
- When a substance is heated, the substance's electrons absorb energy from the flame
- this absorbed energy allows the electrons to move to higher energy levels (they get excited)
- when an electron moves energy levels (the circles) down an energy level, light is being emitted
- the energy being given off (emitted) is in levels and we see this with the color of light

Color that the flame produces

- calcium-
- copper-
- lithium-
- potassium-
- sodium-
- strontium-

Give 3 Examples of how this property can be used in a real life scenario

- 1.
- 2.
- 3.